Molecular, Ionic, and Net Ionic Equations

Write down the molecular form of the following equations. Balance and show ionic and net ionic forms.

1) manganese(II) chloride(aq.) + ammonium carbonate(aq.) → manganese(II) carbonate(s) + ammonium chloride (aq.)

Molecular equation:
Ionic equation:
Net ionic equation:

2) chromium(III) nitrate (aq.) + iron(II) sulfate(aq.) → chromium(III) sulfate(aq.) + iron(II) nitrate(aq.)

Molecular equation:
Ionic equation:
Net ionic equation:

3) strontium bromide(aq.) + potassium sulfate(aq.) → strontium sulfate(s) + potassium bromide(aq.)

Molecular equation:
Ionic equation:
Net ionic equation:

4) copper(II) sulfate(aq.) + potassium carbonate(aq.) → copper(II) carbonate(s) + potassium sulfate(aq.)

Molecular equation:
Ionic equation:
Net ionic equation:

5) cobalt(III) bromide(aq.) + potassium sulfide(aq.) → cobalt sulfide(s) + potassium bromide(aq.)

Molecular equation:
Ionic equation:
Net ionic equation: