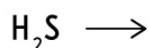
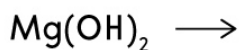
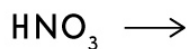


Name : \_\_\_\_\_ Date : \_\_\_\_\_

## SOLUTIONS OF ACIDS AND BASES

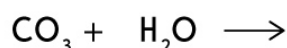
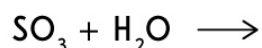
- ① Write the chemical equation showing how each of the following dissolves in water.



- ② For each reaction type, complete the equations and balance them.

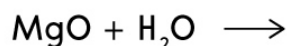
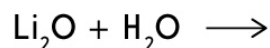
a. Making acids from nonmetal oxide

General form: nonmetal oxide + water  $\longrightarrow$  acid



b. Making bases from metal oxide

General form: Metal oxide + water  $\longrightarrow$  base



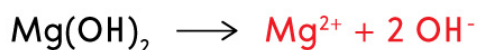
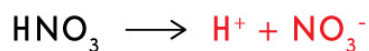
- ③ Write the balanced equation for the titration between sulfurous acid and potassium hydroxide.
- ④ Write the balanced net ionic equation for the titration between sulfurous acid and potassium hydroxide.
- ⑤ Write an equation for the reaction of potassium metal with hydrochloric acid.
- ⑥ Acids can be labelled as monoprotic, diprotic, or triprotic. Label the following as monoprotic, diprotic, or triprotic



# SOLUTIONS OF ACIDS AND BASES

## Answers

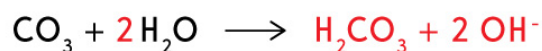
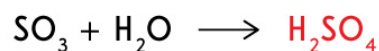
- ① Write the chemical equation showing how each of the following dissolves in water.



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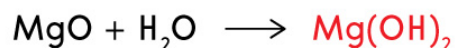
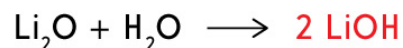
a. Making acids from nonmetal oxide

General form: nonmetal oxide + water  $\longrightarrow$  acid



b. Making bases from metal oxide

General form: Metal oxide + water  $\longrightarrow$  base



- ③ Write the balanced equation for the titration between sulfurous acid and potassium hydroxide.



- ④ Write the balanced net ionic equation for the titration between sulfurous acid and potassium hydroxide.



- ⑤ Write an equation for the reaction of potassium metal with hydrochloric acid.



- ⑥ Acids can be labelled as monoprotic, diprotic, or triprotic. Label the following as monoprotic, diprotic, or triprotic

