

Hess's Law Worksheet

- 1 Determine the heat of formation of liquid hydrogen peroxide at 25°C from the following thermochemical equations.



- 2 Calculate the ΔH for the reaction: $\text{Cu}(\text{s}) + \frac{1}{2} \text{O}_2(\text{g}) \longrightarrow \text{CuO}(\text{s})$

Given the following information:



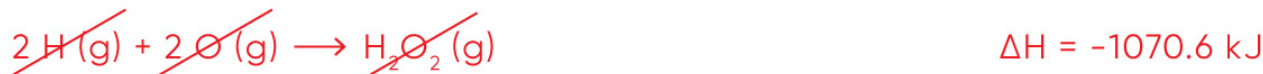
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Answers

- 1 Determine the heat of formation of liquid hydrogen peroxide at 25°C from the following thermochemical equations.



We need to know ΔH for the following reaction: $\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}_2(\text{l})$



- 2 Calculate the ΔH for the reaction: $\text{Cu}(\text{s}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{CuO}(\text{s})$

Given the following information:

