

# Practicing Lewis Dot Structure

Draw the Lewis dot structure of the given molecules.

$N_2$	HF	$H_2S$
HI	$CHCl_3$	$SiH_4$
HCN	$O_2$	$NCl_3$
$CO_3^{2-}$	$SCl_4$	NO
$NH_2Cl$	$CH_3Br$	$CH_2F_2$

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## Answers

$N_2$ $:N \equiv N:$	$HF$ $H - \overset{\cdot\cdot}{\underset{\cdot\cdot}{F}}$	$H_2S$ $H - \overset{\cdot\cdot}{\underset{\cdot\cdot}{S}} - H$
$HI$ $H - \overset{\cdot\cdot}{\underset{\cdot\cdot}{I}}$	$CHCl_3$ $\begin{array}{c} H \\   \\ \overset{\cdot\cdot}{\underset{\cdot\cdot}{Cl}} - C - \overset{\cdot\cdot}{\underset{\cdot\cdot}{Cl}} \\   \\ \overset{\cdot\cdot}{\underset{\cdot\cdot}{Cl}} \end{array}$	$SiH_4$ $\begin{array}{c} H \\   \\ H - Si - H \\   \\ H \end{array}$
$HCN$ $H - C \equiv N:$	$O_2$ $:\overset{\cdot\cdot}{O} = \overset{\cdot\cdot}{O}:$	$NCl_3$ $\begin{array}{c} \overset{\cdot\cdot}{\underset{\cdot\cdot}{Cl}} - \overset{\cdot\cdot}{\underset{\cdot\cdot}{N}} - \overset{\cdot\cdot}{\underset{\cdot\cdot}{Cl}} \\   \\ \overset{\cdot\cdot}{\underset{\cdot\cdot}{Cl}} \end{array}$
$CO_3^{2-}$ $\begin{array}{c} \overset{\cdot\cdot}{\underset{\cdot\cdot}{O}}^- \\   \\ \overset{\cdot\cdot}{\underset{\cdot\cdot}{O}} = C - \overset{\cdot\cdot}{\underset{\cdot\cdot}{O}}^- \end{array}$	$SCl_4$ $\begin{array}{c} \overset{\cdot\cdot}{\underset{\cdot\cdot}{Cl}} \\   \\ \overset{\cdot\cdot}{\underset{\cdot\cdot}{Cl}} - S - \overset{\cdot\cdot}{\underset{\cdot\cdot}{Cl}} \\   \\ \overset{\cdot\cdot}{\underset{\cdot\cdot}{Cl}} \end{array}$	$NO$ $\overset{\cdot\cdot}{N} = \overset{\cdot\cdot}{O}$
$NH_2Cl$ $\begin{array}{c} \overset{\cdot\cdot}{\underset{\cdot\cdot}{Cl}} \\   \\ H - N - H \\ \cdot\cdot \\ \cdot\cdot \end{array}$	$CH_3Br$ $\begin{array}{c} \overset{\cdot\cdot}{\underset{\cdot\cdot}{Br}} \\   \\ H - C - H \\   \\ H \end{array}$	$CH_2F_2$ $\begin{array}{c} H \\   \\ \overset{\cdot\cdot}{\underset{\cdot\cdot}{F}} - C - \overset{\cdot\cdot}{\underset{\cdot\cdot}{F}} \\   \\ H \end{array}$