

# Practicing Lewis Structures

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1. Draw the electron dot structure for the following elements.

Beryllium	Selenium	Tin
Carbon	Sulfur	Iodine

2. Draw the Lewis structure for the following molecules.

$\text{PI}_3$	$\text{AsBr}_3$	$\text{N}_2\text{H}_2$
$\text{CH}_3\text{OH}$	$\text{NO}_2^-$	$\text{FCN}$
$\text{AsN}$	$\text{CH}_3\text{NH}_2$	$\text{HNO}$

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## Answers

Beryllium $\text{Be}\cdot$	Selenium $:\text{Se}\cdot$	Tin $\cdot\text{Sn}\cdot$
Carbon $\cdot\text{C}\cdot$	Sulfur $\cdot\text{S}\cdot$	Iodine $\cdot\text{I}\cdot$

$\text{PI}_3$ 	$\text{AsBr}_3$ 	$\text{N}_2\text{H}_2$ $\text{H}-\ddot{\text{N}}=\ddot{\text{N}}-\text{H}$
$\text{CH}_3\text{OH}$ 	$\text{NO}_2^-$ 	$\text{FCN}$ $:\ddot{\text{F}}-\text{C}\equiv\text{N}:$
$\text{AsN}$ $:\text{As}\equiv\text{N}:$	$\text{CH}_3\text{NH}_2$ 	$\text{HNO}$ $\text{H}-\ddot{\text{N}}=\ddot{\text{O}}$