

Name : \_\_\_\_\_ Date : \_\_\_\_\_

## Solubility Rules

Name or give the chemical formula for each of the following compounds and state whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution.

Chemical Formula	Name	Solubility
$\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$		
$\text{Sr}_2\text{CrO}_4$		
	Aluminum Phosphate	
$\text{BaSO}_4$		
$\text{Ca}(\text{OH})_2$		
$\text{BaCO}_3$		
$\text{MgCrO}_4$		
	Iron (III) sulfide	
$\text{NH}_4\text{CN}$		
	Silver Iodide	
$\text{Hg}_2\text{SO}_4$		
	Silver Iodide	
	Silver Iodide	
$\text{Na}_3\text{PO}_4$		
	Calcium sulfate	
$\text{PbI}_2$		
$\text{Ca}(\text{CN})_2$		
	Cadmium (II) sulfide	
$\text{BeSO}_4$		
	Zinc hydroxide	
$(\text{NH}_4)_2\text{CO}_3$		
$\text{NiCl}_2$		

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## Solubility Rules

### Answers

Name or give the chemical formula for each of the following compounds and state whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution.

Chemical Formula	Name	Solubility
$\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$	Ammonium acetate	Soluble
$\text{Sr}_2\text{CrO}_4$	Strontium chromate	Insoluble
$\text{AlPO}_4$	Aluminum Phosphate	Insoluble
$\text{BaSO}_4$	Barium sulfate	Insoluble
$\text{Ca(OH)}_2$	Calcium hydroxide	Insoluble
$\text{BaCO}_3$	Barium carbonate	Insoluble
$\text{MgCrO}_4$	Magnesium chromate	Insoluble
$\text{Fe}_2\text{S}_3$	Iron (III) sulfide	Insoluble
$\text{NH}_4\text{CN}$	Ammonium cyanide	Soluble
$\text{AgI}$	Silver Iodide	Insoluble
$\text{Hg}_2\text{SO}_4$	Mercury (I) sulfate	Insoluble
$\text{LiCl}$	Silver Iodide	Soluble
$\text{BaBr}_2$	Silver Iodide	Soluble
$\text{Na}_3\text{PO}_4$	Sodium phosphate	Soluble
$\text{CaSO}_4$	Calcium sulfate	Soluble
$\text{PbI}_2$	Lead iodide	Insoluble
$\text{Ca(CN)}_2$	Calcium cyanide	Soluble
$\text{CdS}$	Cadmium (II) sulfide	Insoluble
$\text{BeSO}_4$	Beryllium sulfate	Soluble
$\text{Zn(OH)}_2$	Zinc hydroxide	Insoluble
$(\text{NH}_4)_2\text{CO}_3$	Ammonium carbonate	Soluble
$\text{NiCl}_2$	Nickel chloride	Soluble