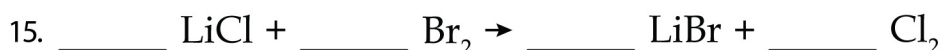
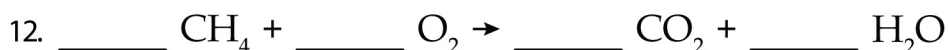
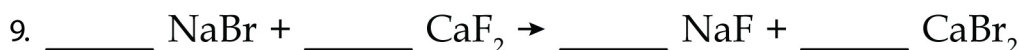
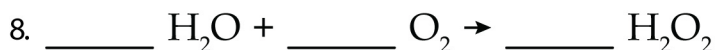
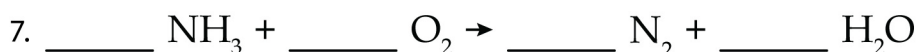
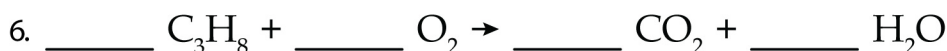
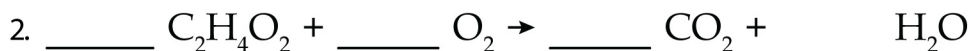


Name : ..... Date : .....

## Balancing Act Worksheet

Balance the following chemical equations.



Name : ..... Date : .....

## Balancing Act Worksheet

### Answers

- $1 \text{ H}_2\text{O} + 1 \text{ SO}_3 \rightarrow 1 \text{ H}_2\text{SO}_4$
- $1 \text{ C}_2\text{H}_4\text{O}_2 + 2 \text{ O}_2 \rightarrow 2 \text{ CO}_2 + 2 \text{ H}_2\text{O}$
- $2 \text{ PbSO}_4 \rightarrow 2 \text{ PbSO}_3 + 1 \text{ O}_2$
- $1 \text{ P}_4 + 6 \text{ Br}_2 \rightarrow 4 \text{ PBr}_3$
- $2 \text{ Mg} + 1 \text{ O}_2 \rightarrow 2 \text{ MgO}$
- $1 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$
- $4 \text{ NH}_3 + 3 \text{ O}_2 \rightarrow 2 \text{ N}_2 + 6 \text{ H}_2\text{O}$
- $2 \text{ H}_2\text{O} + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}_2$
- $2 \text{ NaBr} + 1 \text{ CaF}_2 \rightarrow 2 \text{ NaF} + 1 \text{ CaBr}_2$
- $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$
- $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$
- $1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$
- $1 \text{ AlBr}_3 + 3 \text{ K} \rightarrow 3 \text{ KBr} + 1 \text{ Al}$
- $1 \text{ P}_4 + 6 \text{ Br}_2 \rightarrow 4 \text{ PBr}_3$
- $2 \text{ LiCl} + 1 \text{ Br}_2 \rightarrow 2 \text{ LiBr} + 1 \text{ Cl}_2$