Name : ---------- Date : ----------

Balancing Chemical Equations Practice

Balance the following chemical equations.

(1) _____ Rb + _____ $RbNO_3$ → _____ Rb_2O + _____ N_2

(2) _____ C_6H_6 + _____ O_2 → _____ CO_2 + _____ H_2O

(3) _____ N_2 + _____ H_2 → _____ NH_3

(4) _____ $C_{10}H_{22}$ + _____ O_2 → _____ CO_2 + _____ H_2O

(5) _____ $Al(OH)_3$ + _____ HBr → _____ $AlBr_3$ + _____ H_2O

(6) _____ C_4H_{10} + _____ O_2 → _____ CO_2 + _____ H_2O

(7) _____ C_3H_8 + _____ O_2 → _____ CO_2 + _____ H_2O

(8) _____ Li + _____ $AlCl_3$ → _____ $LiCl$ + _____ Al

(9) _____ C_2H_6 + _____ O_2 → _____ CO_2 + _____ H_2O

(10) _____ NH_4OH + _____ H_3PO_4 → _____ $(NH_4)_3PO_4$ + _____ H_2O

(11) _____ Rb + _____ P → _____ Rb_3P

(12) _____ CH_4 + _____ O_2 → _____ CO_2 + _____ H_2O

(13) _____ $Al(OH)_3$ + _____ H_2SO_4 → _____ $Al_2(SO_4)_3$ + _____ H_2O

(14) _____ Na + _____ Cl_2 → _____ $NaCl$

(15) _____ Rb + _____ S_8 → _____ Rb_2S

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Answers

(1) 10 Rb + 2 $RbNO_3$ → 6 Rb_2O + 1 N_2

(2) 2 C_6H_6 + 15 O_2 → 12 CO_2 + 6 H_2O

(3) 1 N_2 + 3 H_2 → 2 NH_3

(4) 2 $C_{10}H_{22}$ + 31 O_2 → 20 CO_2 + 22 H_2O

(5) 1 $Al(OH)_3$ + 3 HBr → 1 $AlBr_3$ + 3 H_2O

(6) 2 C_4H_{10} + 13 O_2 → 8 CO_2 + 10 H_2O

(7) 1 C_3H_8 + 5 O_2 → 3 CO_2 + 4 H_2O

(8) 3 Li + 1 $AlCl_3$ → 3 $LiCl$ + 1 Al

(9) 2 C_2H_6 + 7 O_2 → 4 CO_2 + 6 H_2O

(10) 3 NH_4OH + 1 H_3PO_4 → 1 $(NH_4)_3PO_4$ + 3 H_2O

(11) 3 Rb + 1 P → 1 Rb_3P

(12) 1 CH_4 + 2 O_2 → 1 CO_2 + 2 H_2O

(13) 2 $Al(OH)_3$ + 3 H_2SO_4 → 1 $Al_2(SO_4)_3$ + 6 H_2O

(14) 2 Na + 1 Cl_2 → 2 $NaCl$

(15) 16 Rb + 1 S_8 → 8 Rb_2S