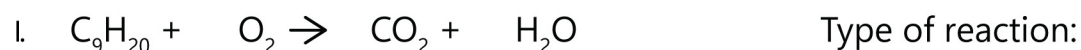
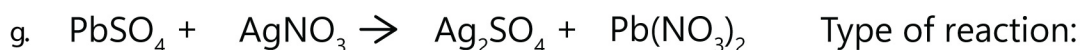
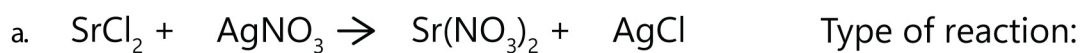


Name : _____ Date : _____

Balancing and Classifying Chemical Equations

Balance the chemical reactions and identify each type.



Name : _____ Date : _____

Balancing and Classifying Chemical Equations

Answers

a. $1 \text{ SrCl}_2 + 2 \text{ AgNO}_3 \rightarrow 1 \text{ Sr(NO}_3)_2 + 2 \text{ AgCl}$ Type of reaction: Double displacement

b. $2 \text{ LiBr} + 1 \text{ F}_2 \rightarrow 2 \text{ LiF} + 1 \text{ Br}_2$ Type of reaction: Single displacement

c. $2 \text{ NaF} \rightarrow 2 \text{ Na} + 1 \text{ F}_2$ Type of reaction: Decomposition

d. $1 \text{ ZnBr}_2 + 1 \text{ F}_2 \rightarrow 1 \text{ ZnF}_2 + 1 \text{ Br}_2$ Type of reaction: Single displacement

e. $1 \text{ CuCl}_2 + 1 \text{ H}_2\text{S} \rightarrow 1 \text{ CuS} + 2 \text{ HCl}$ Type of reaction: Double displacement

f. $1 \text{ Ca} + 2 \text{ AgCl} \rightarrow 2 \text{ CaCl}_2 + 2 \text{ Ag}$ Type of reaction: Single displacement

g. $1 \text{ PbSO}_4 + 2 \text{ AgNO}_3 \rightarrow 1 \text{ Ag}_2\text{SO}_4 + 1 \text{ Pb(NO}_3)_2$ Type of reaction: Double displacement

h. $6 \text{ Li} + 1 \text{ N}_2 \rightarrow 2 \text{ Li}_3\text{N}$ Type of reaction: Synthesis

i. $1 \text{ C}_4\text{H}_8 + 6 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 4 \text{ H}_2\text{O}$ Type of reaction: Combustion

j. $3 \text{ Ca(OH)}_2 + 1 \text{ Al}_2(\text{SO}_4)_3 \rightarrow 3 \text{ CaSO}_4 + 2 \text{ Al(OH)}_3$ Type of reaction: Double displacement

k. $2 \text{ NH}_3 + 3 \text{ I}_2 \rightarrow 1 \text{ N}_2\text{I}_6 + 3 \text{ H}_2$ Type of reaction: Single displacement

l. $1 \text{ C}_9\text{H}_{20} + 14 \text{ O}_2 \rightarrow 9 \text{ CO}_2 + 10 \text{ H}_2\text{O}$ Type of reaction: Combustion