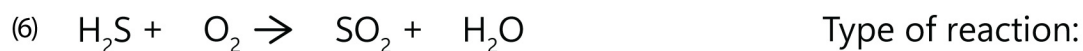
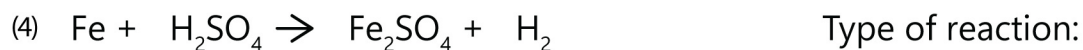
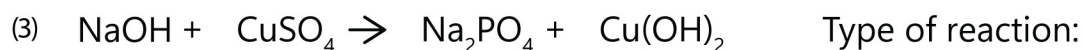
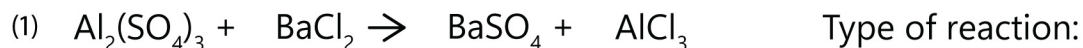


Name : _____ Date : _____

BALANCING AND IDENTIFYING CHEMICAL EQUATIONS

Classify the following reactions and balance them.



Name : _____ Date : _____

BALANCING AND IDENTIFYING CHEMICAL EQUATIONS

Answers

- (1) $1 \text{ Al}_2(\text{SO}_4)_3 + 3 \text{ BaCl}_2 \rightarrow 3 \text{ BaSO}_4 + 2 \text{ AlCl}_3$ Type of reaction: Double displacement
- (2) $1 \text{ Al}_2\text{S}_3 \rightarrow 2 \text{ Al} + 3 \text{ S}$ Type of reaction: Decomposition
- (3) $2 \text{ NaOH} + 1 \text{ CuSO}_4 \rightarrow 1 \text{ Na}_2\text{PO}_4 + 1 \text{ Cu}(\text{OH})_2$ Type of reaction: Double displacement
- (4) $1 \text{ Fe} + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ Fe}_2\text{SO}_4 + 1 \text{ H}_2$ Type of reaction: Single displacement
- (5) $1 \text{ C}_4\text{H}_{12} + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$ Type of reaction: Combustion
- (6) $2 \text{ H}_2\text{S} + 3 \text{ O}_2 \rightarrow 2 \text{ SO}_2 + 2 \text{ H}_2\text{O}$ Type of reaction: Redox
- (7) $4 \text{ C}_5\text{H}_9\text{O} + 27 \text{ O}_2 \rightarrow 20 \text{ CO}_2 + 18 \text{ H}_2\text{O}$ Type of reaction: Combustion
- (8) $2 \text{ Al} + 3 \text{ NiBr}_2 \rightarrow 2 \text{ AlBr}_3 + 3 \text{ Ni}$ Type of reaction: Single displacement
- (9) $4 \text{ Al} + 3 \text{ O}_2 \rightarrow 2 \text{ Al}_2\text{O}_3$ Type of reaction: Synthesis
- (10) $2 \text{ H}_2\text{O}_2 \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ O}_2$ Type of reaction: Decomposition
- (11) $2 \text{ K} + 1 \text{ Cl}_2 \rightarrow 2 \text{ KCl}$ Type of reaction: Synthesis
- (12) $2 \text{ Na} + 1 \text{ MgCl}_2 \rightarrow 2 \text{ NaCl} + 1 \text{ Mg}$ Type of reaction: Single displacement