Name : ._____

----- Date : -----

Types of Reactions Worksheet

First, begin by writing which type of reaction is occurring. Then, go back and balance the following equations:

1
$$Zn + HCl \rightarrow ZnCl_2 + H_2$$

Type of reaction:

Type of reaction:

3 NaBr +
$$H_3PO_4 \rightarrow Na_3PO_4 + HBr$$

Type of reaction:

4
$$Ca(OH)_2 + Al_2(SO_4)_3 \rightarrow CaSO_4 + Al(OH)_3$$
 Type of reaction:

5 Mg +
$$Fe_2O_3 \rightarrow Fe + MgO$$

Type of reaction:

Type of reaction:

7
$$PbSO_4 \rightarrow PbSO_3 + O_2$$

Type of reaction:

8
$$NH_3 + I_2 \rightarrow N_2I_6 + H_2$$

Type of reaction:

$$9 \quad H_2O + SO_3 \rightarrow H_2SO_4$$

Type of reaction:

10
$$H_2SO_4 + NH_4OH \rightarrow H_2O + (NH_4)_2SO_4$$

Type of reaction:

11
$$H_2SO_4 + Fe \rightarrow H_2 + FeSO_4$$

Type of reaction:

Type of reaction:

Types of Reactions Worksheet

Answers

1
$$1 \text{ Zn} + 2 \text{ HCl} \rightarrow 1 \text{ ZnCl}_2 + 1 \text{ H}_2$$

Type of reaction: Single displacement

$$2 + P + 5 O_2 \rightarrow 1 P_4 O_{10}$$

Type of reaction: Synthesis

3 NaBr +
$$1 H_3 PO_4 \rightarrow 1 Na_3 PO_4 + 3 HBr$$

Type of reaction: Double displacement

4 3 Ca(OH)₂ + 1 Al₂(SO₄)₃
$$\rightarrow$$
 3 CaSO₄ + 2 Al(OH)₃ Type of reaction: Double displacement

5
 3 Mg + 1 Fe₂O₃ \rightarrow 2 Fe + 3 MgO

Type of reaction: Single displacement

$$6 1 C_2 H_4 + 3 O_2 \rightarrow 2 CO_2 + 2 H_2 O$$

Type of reaction: Combustiont

7 2 PbSO₄
$$\rightarrow$$
 2 PbSO₃ + 1 O₂

Type of reaction: Decomposition

$$|8| 2 NH_3 + 3 I_2 \rightarrow 1 N_2 I_6 + 3 H_2$$

Type of reaction: Single displacement

$$9 1 H_2O + 1 SO_3 \rightarrow 1 H_2SO_4$$

Type of reaction: Combination

10
$$1 H_2 SO_4 + 2 NH_4 OH \rightarrow 2 H_2 O + 1 (NH_4)_2 SO_4$$

Type of reaction: Acid-base

11
$$1 H_2 SO_4 + 1 Fe \rightarrow 1 H_2 + 1 Fe SO_4$$

Type of reaction: Single displacement

12
$$\frac{1}{1}$$
 HCl + $\frac{1}{1}$ NH₃ $\rightarrow \frac{1}{2}$ NH₄Cl

Type of reaction: Synthesis