

Name : _____ Date : _____

Oxidation Number of Ions Worksheet

What is the oxidation number of each element in the following compounds?

Compound	Oxidation Numbers of Each Element			
Ag_2S				
Al_2O_3				
Ca_3P_2				
ClO_3^-				
HI				
HNO_2				
H_3O^+				
KMnO_4				
N_2H_4				
$(\text{NH}_4)_2\text{SO}_4$				
O_2				
S_8				
$\text{S}_2\text{O}_3^{2-}$				
SnCl_4				

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Answers

Compound	Oxidation Numbers of Each Element			
Ag_2S	Ag = +1	S = -2		
Al_2O_3	Al = +3	O = -2		
Ca_3P_2	Ca = +2	P = -3		
ClO_3^-	Cl = +5	O = -2		
HI	H = +1	I = -1		
HNO_2	H = +1	N = +3	O = -2	
H_3O^+	H = +1	O = -2		
KMnO_4	K = +1	Mn = +7	O = -2	
N_2H_4	N = -2	H = +1		
$(\text{NH}_4)_2\text{SO}_4$	N = -3	H = +1	S = +6	O = -2
O_2	O = 0			
S_8	S = 0			
$\text{S}_2\text{O}_3^{2-}$	S = +2	O = -2		
SnCl_4	Sn = +4	Cl = -1		