

Name : _____ Date : _____

Oxidation Numbers MCQ Worksheet

Circle the correct option:

- Which of the following is the correct sum of the oxidation numbers of the nitrate (NO_3^-) ion?
 - $\text{N} + 3\text{O} = 0$
 - $\text{N} + 3\text{O} = -1$
 - $\text{N} + 3\text{O} = 1$
 - $\text{N} + \text{O} = 3$
- What is the oxidation number of sulfur in sulphuric acid (H_2SO_4)?
 - 4
 - 6
 - 8
 - 2
- The oxidation numbers of N_2 add up to
 - $2\text{N} = 0$
 - $\text{N} = 0$
 - $2\text{N} = -3$
 - $\text{N} = 3$
- Which of the following statements between HClO_4 and HClO_3 is true?
 - The oxidation number of Cl in HClO_4 has decreased in HClO_3
 - The oxidation numbers for all atoms have remained the same in both molecules
 - The oxidation number for Cl in HClO_4 has increased in HClO_3
 - The oxidation number of O in HClO_4 has decreased in HClO_3
- In the phosphate (PO_4^{3-}) ion, which of the following is a correct representation of the sum of the oxidation numbers of its individual elements?
 - $\text{P} + 4\text{O} = -3$
 - $3 \times [\text{P} + 4\text{O}] = 3$
 - $\text{P} + 4\text{O} = 3$
 - $\text{P} + 4\text{O} = 0$

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Answers

- Which of the following is the correct sum of the oxidation numbers of the nitrate (NO_3^-) ion?
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 - $\text{P} + 4\text{O} = 3$
 - $\text{P} + 4\text{O} = 0$