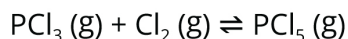


Name : Date :

Worksheet on Le Chatelier's Principle

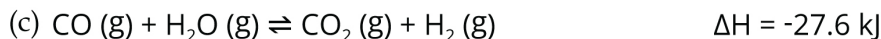
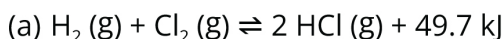
1. For the reaction



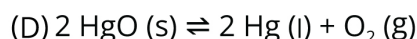
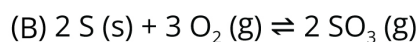
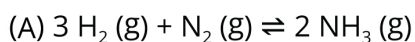
Predict the effect on the position of equilibrium that results from :

Stress	Shift	Stress	Shift
Decreasing the volume		Addition of Ne, an inert gas	
Removing some PCl_5		Increasing the temperature	
Injecting more Cl_2 gas		Increasing the volume of the container	

2. Predict the effect of decreasing the temperature on the position of the following equilibria.

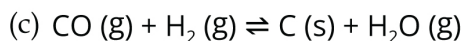
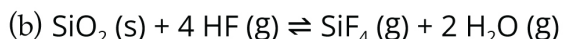
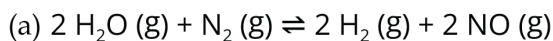


3. Which system at equilibrium will be least affected by a change in pressure ?



Ans: _____

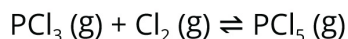
4. Predict the effect of decreasing the container volume for each equilibrium.



Name : _____ Date : _____

Worksheet on Le Chatelier's Principle

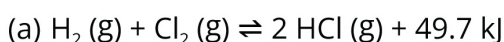
1. For the reaction



Predict the effect on the position of equilibrium that results from :

Stress	Shift	Stress	Shift
Decreasing the volume	Right	Addition of Ne, an inert gas	None
Removing some PCl_5	Right	Increasing the temperature	Left
Injecting more Cl_2 gas	Right	Increasing the volume of the container	Left

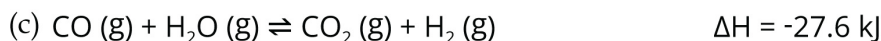
2. Predict the effect of decreasing the temperature on the position of the following equilibria.



Equilibrium will shift to favor products.

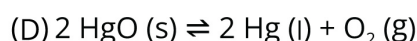
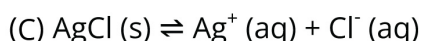
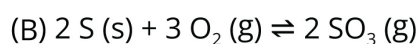
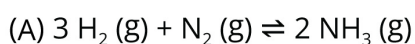


Equilibrium will shift to favor reactants.



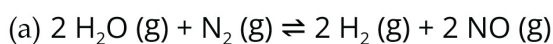
Equilibrium will shift to favor products.

3. Which system at equilibrium will be least affected by a change in pressure ?

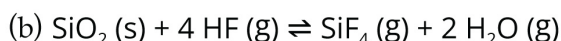


Ans: C

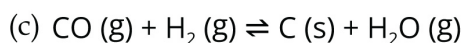
4. Predict the effect of decreasing the container volume for each equilibrium.



Shifts left to produce fewer number of gas molecules.



Shifts right to produce fewer number of moles of gas.



Shifts right to produce fewer number of moles of gas.