

# Ions and Ionic Compounds Worksheet

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Fill in the table with the correct information.

<b>Name of the Compound</b>	<b>Positive Ion</b>	<b>Negative Ion</b>	<b>Formula</b>
Sodium Chloride			
Zinc Chloride			
Calcium Sulphide			
Potassium Bromide			
Lithium Oxide			
Calcium Oxide			
Potassium Nitride			
Sodium Sulphide			
Aluminum Chloride			
Sodium Phosphide			
Magnesium Phosphide			
Aluminum Phosphide			
Strontium Fluoride			
Aluminum Oxide			
Calcium Nitride			
Magnesium Sulphide			
Aluminum Sulphide			
Zinc Nitride			
Silver Sulphide			
Zinc Sulphide			

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## Answers

Name of the Compound	Positive Ion	Negative Ion	Formula
Sodium Chloride	$\text{Na}^+$	$\text{Cl}^-$	$\text{NaCl}$
Zinc Chloride	$\text{Zn}^{2+}$	$\text{Cl}^-$	$\text{ZnCl}_2$
Calcium Sulphide	$\text{Ca}^{2+}$	$\text{S}^{2-}$	$\text{CaS}$
Potassium Bromide	$\text{K}^+$	$\text{Br}^-$	$\text{KBr}$
Lithium Oxide	$\text{Li}^+$	$\text{O}^{2-}$	$\text{Li}_2\text{O}$
Calcium Oxide	$\text{Ca}^{2+}$	$\text{O}^{2-}$	$\text{CaO}$
Potassium Nitride	$\text{K}^+$	$\text{N}^{3-}$	$\text{K}_3\text{N}$
Sodium Sulphide	$\text{Na}^+$	$\text{S}^{2-}$	$\text{Na}_2\text{S}$
Aluminum Chloride	$\text{Al}^{3+}$	$\text{Cl}^-$	$\text{AlCl}_3$
Sodium Phosphide	$\text{Na}^+$	$\text{P}^{3-}$	$\text{Na}_3\text{P}$
Magnesium Phosphide	$\text{Mg}^{2+}$	$\text{P}^{3-}$	$\text{Mg}_3\text{P}_2$
Aluminum Phosphide	$\text{Al}^{3+}$	$\text{P}^{3-}$	$\text{AlP}$
Strontium Fluoride	$\text{Sr}^{2+}$	$\text{F}^-$	$\text{SrF}_2$
Aluminum Oxide	$\text{Al}^{3+}$	$\text{O}^{2-}$	$\text{Al}_2\text{O}_3$
Calcium Nitride	$\text{Ca}^{2+}$	$\text{N}^{3-}$	$\text{Ca}_3\text{N}_2$
Magnesium Sulphide	$\text{Mg}^{2+}$	$\text{S}^{2-}$	$\text{MgS}$
Aluminum Sulphide	$\text{Al}^{3+}$	$\text{S}^{2-}$	$\text{Al}_2\text{S}_3$
Zinc Nitride	$\text{Zn}^{2+}$	$\text{N}^{3-}$	$\text{Zn}_3\text{N}_2$
Silver Sulphide	$\text{Ag}^+$	$\text{S}^{2-}$	$\text{Ag}_2\text{S}$
Zinc Sulphide	$\text{Zn}^{2+}$	$\text{S}^{2-}$	$\text{ZnS}$