

# Naming Compounds Worksheet

1. Name the following ionic compounds.

a) NaOH \_\_\_\_\_

b) FeCl<sub>2</sub> \_\_\_\_\_

c) FeCl<sub>3</sub> \_\_\_\_\_

d) Zn(OH)<sub>2</sub> \_\_\_\_\_

e) PbO \_\_\_\_\_

2. Name the following covalent compounds.

a) CH<sub>4</sub> \_\_\_\_\_

b) NF<sub>3</sub> \_\_\_\_\_

c) PH<sub>3</sub> \_\_\_\_\_

d) NO<sub>2</sub> \_\_\_\_\_

e) SF<sub>6</sub> \_\_\_\_\_

3. Identify the following compounds and mention whether they are ionic or covalent.

a) BF<sub>3</sub> \_\_\_\_\_

b) MgBr<sub>2</sub> \_\_\_\_\_

c) NH<sub>4</sub>OH \_\_\_\_\_

d) HI \_\_\_\_\_

e) N<sub>2</sub>O<sub>3</sub> \_\_\_\_\_

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1.

Answers

a) NaOH Sodium Hydroxide

b) FeCl<sub>2</sub> Ferrous Chloride or Iron(II) Chloride

c) FeCl<sub>3</sub> Ferric Chloride or Iron(III) Chloride

d) Zn(OH)<sub>2</sub> Zinc Hydroxide

e) PbO Lead(II) Oxide

2.

a) CH<sub>4</sub> Methane

b) NF<sub>3</sub> Nitrogen Trifluoride

c) PH<sub>3</sub> Phosphorus Trihydride

d) NO<sub>2</sub> Nitrogen Dioxide

e) SF<sub>6</sub> Sulphur Hexafluoride

3.

a) BF<sub>3</sub> Boron Trifluoride; Covalent Compound

b) MgBr<sub>2</sub> Magnesium Dibromide; Ionic Compound

c) NH<sub>4</sub>OH Ammonium Hydroxide; Ionic Compound

d) HI Hydrogen Iodide; Covalent Compound

e) N<sub>2</sub>O<sub>3</sub> Dinitrogen Trioxide; Covalent Compound