

Name : \_\_\_\_\_ Date : \_\_\_\_\_

# Ionic Bonding Worksheet

Fill in the blanks with the correct information.

1. The bonds formed between \_\_\_\_\_ and \_\_\_\_\_ are called ionic bonds.
2. When naming simple ionic compounds, the metal is always the \_\_\_\_\_ part of the compound, while the non-metal is the \_\_\_\_\_ part.
3. Ionic compounds tend to be \_\_\_\_\_ in water and other polar solvents.
4. Ionic compounds tend to be \_\_\_\_\_ conductors of electricity when in the form of a solution.
5. Ionic compounds tend to be \_\_\_\_\_ solids.
6. Ionic compounds tend to melt at \_\_\_\_\_ temperatures.
7. \_\_\_\_\_ is formed when the metal sodium reacts with the non-metal chlorine.
8. In an ionic compound, the \_\_\_\_\_ tells the ratio of the atoms present in it.
9. \_\_\_\_\_ is made up of  $\text{Fe}^{3+}$  and  $3 \text{Cl}^-$  ions.
10. When two negative ions end up close, they end up \_\_\_\_\_ each other.
11. When positive and negative ions end up close, they \_\_\_\_\_ each other.
12. Metals tend to be \_\_\_\_\_ ions, while non-metals tend to be \_\_\_\_\_ ions.

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# Ionic Bonding Worksheet

Fill in the blanks with the correct information.

1. The bonds formed between metals and non-metals are called ionic bonds.
2. When naming simple ionic compounds, the metal is always the first part of the compound, while the non-metal is the second part.
3. Ionic compounds tend to be soluble in water and other polar solvents.
4. Ionic compounds tend to be good conductors of electricity when in the form of a solution.
5. Ionic compounds tend to be crystalline solids.
6. Ionic compounds tend to melt at high temperatures.
7. Sodium chloride is formed when the metal sodium reacts with the non-metal chlorine.
8. In an ionic compound, the chemical formula tells the ratio of the atoms present in it.
9. FeCl<sub>3</sub> is made up of Fe<sup>3+</sup> and 3 Cl<sup>-</sup> ions.
10. When two negative ions end up close, they end up repelling each other.
11. When positive and negative ions end up close, they attract each other.
12. Metals tend to be positive ions, while non-metals tend to be negative ions.