

Name : \_\_\_\_\_ Date : \_\_\_\_\_

## Ionic Bonds Fill-in-the-Blanks Worksheet

Fill in the blanks with the correct information.

1. Ionic bonds \_\_\_\_\_ the electrons, while a covalent bond \_\_\_\_\_ the electrons.
2. The \_\_\_\_\_ in electronegativity determines the type of bond formed.
3. Ionic compounds easily conduct \_\_\_\_\_ in their liquid state.
4. Ionic compounds tend to melt at \_\_\_\_\_ temperatures.
5. Generally, a metal \_\_\_\_\_ electrons from its outermost shell to a \_\_\_\_\_ in an ionic bond.
6. A positively charged ion is called a \_\_\_\_\_ .
7. An \_\_\_\_\_ is a negatively charged ion.
8. \_\_\_\_\_ is the most electronegative element.
9. The energy required to \_\_\_\_\_ the ionic bonds within a crystal lattice is called lattice energy.
10. A polyatomic ion is one made up of several \_\_\_\_\_ and has an overall \_\_\_\_\_ .
11. Only \_\_\_\_\_ electrons take part in chemical bonding.
12. Ionic bonds are the \_\_\_\_\_ of all the chemical bonds.
13. The formation of cations, i.e., the transfer of an electron from a metal is an \_\_\_\_\_ process.
14. Purely ionic bonding \_\_\_\_\_ exist.
15. \_\_\_\_\_ bonds display partial ionic and covalent characteristics.

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## Ionic Bonds Fill-in-the-Blanks Worksheet

Fill in the blanks with the correct information.

1. Ionic bonds transfer the electrons, while a covalent bond shares the electrons.
2. The difference in electronegativity determines the type of bond formed.
3. Ionic compounds easily conduct electricity in their liquid state.
4. Ionic compounds tend to melt at high temperatures.
5. Generally, a metal transfers electrons from its outermost shell to a non-metal in an ionic bond.
6. A positively charged ion is called a cation .
7. An anion is a negatively charged ion.
8. Fluorine is the most electronegative element.
9. The energy required to break the ionic bonds within a crystal lattice is called lattice energy.
10. A polyatomic ion is one made up of several elements and has an overall charge .
11. Only valence electrons take part in chemical bonding.
12. Ionic bonds are the strongest of all the chemical bonds.
13. The formation of cations, i.e., the transfer of an electron from a metal is an endothermic process.
14. Purely ionic bonding doesn't exist.
15. Polar covalent bonds display partial ionic and covalent characteristics.