

# IONIC AND COVALENT COMPOUNDS WORKSHEET



1. What are the different features of ionic and covalent compounds? (Answer in the form of a table).

Ans:

2. Give a few examples of both ionic and covalent compounds.

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## Answers

1. What are the different features of ionic and covalent compounds? (Answer in the form of a table).

Ans:

Ionic Compounds	Covalent Compounds
These tend to be in the form of crystalline solids.	These can be in the form of solids, liquids, or gases.
Ionic compounds tend to be formed between metals and non-metals.	Covalent compounds tend to be formed between non-metals.
A stable molecule is formed when a metal transfers its electrons to a non-metal.	Stability is achieved through the sharing of electrons between non-metals.
High melting and boiling points.	Low melting and boiling points.
Capable of conducting electricity, especially in a liquid state.	Poor conductors in all states of matter.
Tend to be water-soluble.	Tend to be insoluble in water.
Tends to be insoluble in non-polar liquids.	Tends to be soluble in polar liquids.

2. Give a few examples of both ionic and covalent compounds.

Examples of ionic compounds: NaCl, KF, MgCl<sub>2</sub>, etc.

Examples of covalent compounds: CO, H<sub>2</sub>O, NH<sub>3</sub>, etc.

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