

Name : Date :

Metallic Bonding True or False Worksheet

Are the following statements true or false?

1. Metallic bonds take place between metallic cations and delocalized electrons →
2. Metals are extremely brittle →
3. Metals are great conductors of electricity →
4. Metals are poor conductors of heat →
5. Metals are dull to look at →
6. Metals are malleable thanks to their rigid structural lattice →
7. Metals are ductile due to their flexible structure →
8. In a metal lattice, the metal ions form ionic bonds with electrons from the "sea of delocalized electrons" →
9. Metals can only be found in Group 1A of the periodic table →
10. The smaller the metallic radius, the stronger the metallic bond →
11. Magnesium has 3 delocalized electrons per ion →
12. The charge of the metallic ion plays a key role in the type of metallic bond it forms →
13. Steel is an alloy of iron and copper →
14. Sterling silver is an alloy of silver and copper →
15. The fewer delocalized electrons, the better the metallic bond →

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Answers

Are the following statements true or false?

1. Metallic bonds take place between metallic cations and delocalized electrons → **True**
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14. Sterling silver is an alloy of silver and copper → **True**
15. The fewer delocalized electrons, the better the metallic bond → **False**