Polyatomic Ion Practice Worksheet

Write the formulae for the following polyatomic ionic compounds.

1. Iron (III) Bromate _____
2. Tin (IV) Acetate _____
3. Copper (I) Sulfate _____
4. Arsenic (V) Hypophosphite _____
5. Barium Iodate _____
6. Copper (I) Persulfate _____
7. Manganese (II) Carbonate _____
8. Antimony (III) Perphosphate _____
9. Nickel (III) Manganate _____
10. Zinc Bromate _____
11. Nickel (III) Permanganate _____
12. Gold (I) Phosphite _____
13. Mercury (I) Sulfate _____
14. Potassium Cyanate _____
15. Iron (III) Thiosulfate _____
16. Mercury (I) Bicarbonate _____
17. Rubidium Perchlorate _____
18. Aluminum Acetate _____
19. Silver Sulfite _____
20. Magnesium Hypocarbonite _____

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Answers

1. Iron (III) Bromate — $Fe(BrO_3)_3$
2. Tin (IV) Acetate — $Sn(C_2H_3O_2)_4$
3. Copper (I) Sulfate — Cu_2SO_4
4. Arsenic (V) Hypophosphite — $As_3(PO_2)_5$
5. Barium Iodate — $Ba(IO_3)_2$
6. Copper (I) Persulfate — Cu_2SO_5
7. Manganese (II) Carbonate — $MnCO_3$
8. Antimony (III) Perphosphate — $SbPO_5$
9. Nickel (III) Manganate — $Ni(MnO_3)_3$
10. Zinc Bromate — $Zn(BrO_3)_2$
11. Nickel (III) Permanganate — $Ni(MnO_4)_3$
12. Gold (I) Phosphite — Au_3PO_3
13. Mercury (I) Sulfate — Hg_2SO_4
14. Potassium Cyanate — $KCNO$
15. Iron (III) Thiosulfate — $Fe_2(S_2O_3)_3$
16. Mercury (I) Bicarbonate — $HgHCO_3$
17. Rubidium Perchlorate — $RbClO_2$
18. Aluminum Acetate — $Al(C_2H_3O_2)_3$
19. Silver Sulfite — Ag_2SO_3
20. Magnesium Hypocarbonite — $MgCO$