

Name : _____ Date : _____

Atomic Number and Mass Number Practice Worksheet

1. What is the difference between atomic number and mass number?

2. Complete the table with the correct information.

Atomic Number	Mass Number	Element
9		
		Silicon
22	47	
	55	
		Bromine
8	16	
	108	
16		
		Lead
79		

3. How many subatomic particles (protons, electrons, and neutrons) are in the following atom?

Beryllium (Be)

Atomic Number = 4

Mass Number = 9

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Answers

1. What is the difference between atomic number and mass number?

The key difference between atomic number and mass number is that the former is the number of protons present in an atom, while the latter is the number of protons and neutrons combined.

2. Complete the table with the correct information.

Atomic Number	Mass Number	Element
9	19	Fluorine
14	29	Silicon
22	47	Titanium
25	55	Manganese
35	80	Bromine
8	16	Oxygen
47	108	Silver
16	32	Sulfur
82	207	Lead
79	197	Gold

3. How many subatomic particles (protons, electrons, and neutrons) are in the following atom?

Beryllium (Be)

Atomic Number = 4

Mass Number = 9

Number of Protons = Number of Electrons = 4

Number of Neutrons = Mass Number - Atomic Number = $9 - 4 = 5$