

Electrons in Atoms Worksheet

How many atoms are in the following compounds? (Only number of atoms)

a) Hydrogen

b) Sulfur

c) Oxygen

d) Magnesium

e) Arsenic

f) Chlorine

g) Nitrogen

h) Bromine

i) Vanadium

j) Neon

k) Radon

l) Calcium

m) Iron

n) Fluorine

o) Barium

p) Chromium

q) Phosphorus

r) Copper

s) Scandium

t) Boron

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Answers

a) Hydrogen	$1s^1$
b) Sulfur	$[\text{Ne}] 3s^2 3p^4$
c) Oxygen	$[\text{He}] 2s^2 2p^4$
d) Magnesium	$[\text{Ne}] 3s^2$
e) Arsenic	$[\text{Ar}] 3d^{10} 4s^2 4p^3$
f) Chlorine	$[\text{Ne}] 3s^2 3p^5$
g) Nitrogen	$[\text{He}] 2s^2 2p^3$
h) Bromine	$[\text{Ar}] 3d^{10} 4s^2 4p^5$
i) Vanadium	$[\text{Ar}] 3d^3 4s^2$
j) Neon	$[\text{He}] 2s^2 2p^6$
k) Radon	$[\text{Xe}] 4f^{14} 5d^{10} 6s^2 6p^6$
l) Calcium	$[\text{Ar}] 4s^2$
m) Iron	$[\text{Ar}] 3d^6 4s^2$
n) Fluorine	$[\text{He}] 2s^2 2p^5$
o) Barium	$[\text{Xe}] 6s^2$
p) Chromium	$[\text{Ar}] 3d^5 4s^1$
q) Phosphorus	$[\text{Ne}] 3s^2 3p^3$
r) Copper	$[\text{Ar}] 3d^{10} 4s^1$
s) Scandium	$[\text{Ar}] 4s^2 3d^1$
t) Boron	$[\text{He}] 2s^2 2p^1$