GRAM FORMULA
MASS WORKSHEET

Determine the gram formula mass of the following compounds.
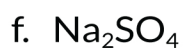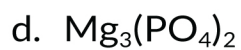
a. H_2CO_3

b. KCl

c. $Fe_2(SO_3)_3$

d. $Mg_3(PO_4)_2$

e. $KMnO_4$

f. Na_2SO_4

g. $Ba(ClO_3)_2$

h. $(NH_4)_3PO_4$

i. $Hg_2Cr_2O_7$

j. $(NH_4)C_2H_3O_2$
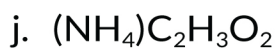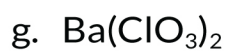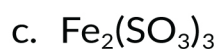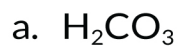
Name : _____

Date : _____

GRAM FORMULA
MASS WORKSHEET

a. H_2CO_3
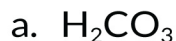
Gram Formula Mass of H_2CO_3 = (2 x 1) + (1 x 12) + (3 x 16) = 62 g

b. KCl

Gram Formula Mass of KCl = (1 x 39) + (1 x 35.5) = 74.5 g
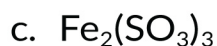
c. $Fe_2(SO_3)_3$

Gram Formula Mass of $Fe_2(SO_3)_3$ = (2 x 55.845) + (3 x 32) + (9 x 16) = 351.69 g
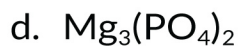
d. $Mg_3(PO_4)_2$

Gram Formula Mass of $Mg_3(PO_4)_2$ = (3 x 24) + (2 x 30.97) + (8 x 16) = 261.94 g

e. $KMnO_4$

Gram Formula Mass of $KMnO_4$ = (1 x 39) + (1 x 54.9) + (4 x 16) = 157.9 g

f. Na_2SO_4

Gram Formula Mass of Na_2SO_4 = (2 x 23) + (1 x 32) + (4 x 16) = 142 g

g. $Ba(ClO_3)_2$

Gram Formula Mass of $Ba(ClO_3)_2$ = (1 x 137.3) + (2 x 35.5) + (6 x 16) = 304.3 g

h. $(NH_4)_3PO_4$

Gram Formula Mass of $(NH_4)_3PO_4$ = (3 x 14) + (7 x 1) + (1 x 30.97) + (4 x 16) = 143.97 g

i. $Hg_2Cr_2O_7$

Gram Formula Mass of $Hg_2Cr_2O_7$ = (2 x 200.5) + (2 x 52) + (7 x 16) = 617 g
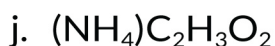
j. $(NH_4)C_2H_3O_2$

Gram Formula Mass of $(NH_4)C_2H_3O_2$ = (1 x 14) + (7 x 1) + (2 x 12) + (2 x 16) = 77 g

Name : _____

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